Stoichiometry Problem Guidelines:

1. _____________ the equation

2. convert to _____________ of _____________ substance

3. MAKE THE _____________ USING _____________ _____________ FROM BALANCED EQUATION

4. convert to _____________

Problem Set One
Limiting Reactant

____________ used up __________ in a ______________ reaction.

Excess Reactant

____________ that is not used up in a chemical ______________

Ex. Problem:
When FeCl₃ reacts with O₂, Fe₂O₃ and Cl₂ are produced. If 4.0 moles of FeCl₃ and 4.0 moles of O₂ are mixed, how many grams of Fe₂O₃ will be produced?

→

(Hint: Work two separate problems, using one reactant at a time.)

(Hint: Answer will be the (smaller, larger) amount of product.)

What is the limiting reactant?

What is the excess reactant?

Problem Set Two (Work on back.)

The Chemistry Quiz

CR1.           CR2.       1.       2.

3.               4.           5.

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8.28